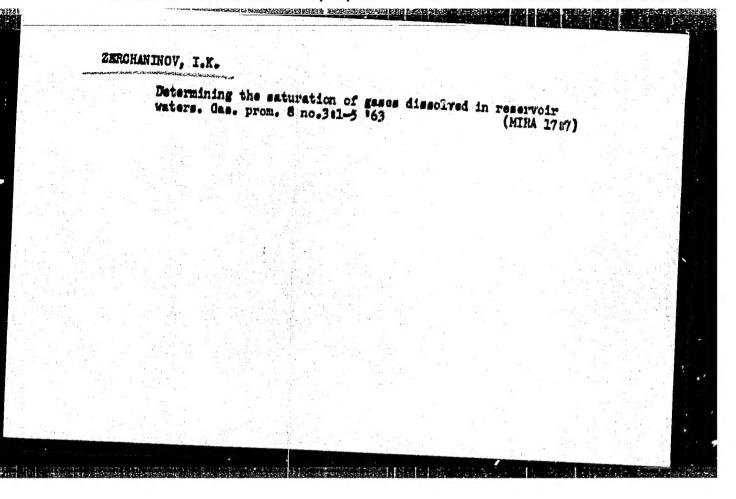
IMOV, I.K.			
Nethod for calculat VNII no.30:128-135	ing reduced pressure '60. (Oil field brines)	s of formation	waters. Truly (MIMA 14:2)

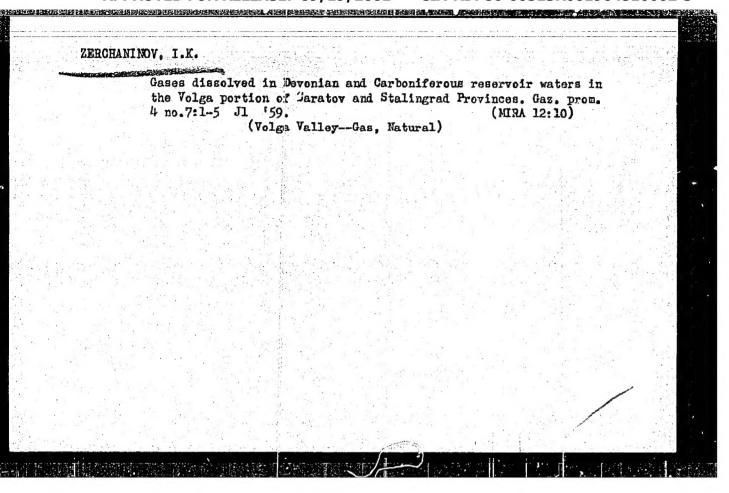
ZERCHANINOV, I.K.

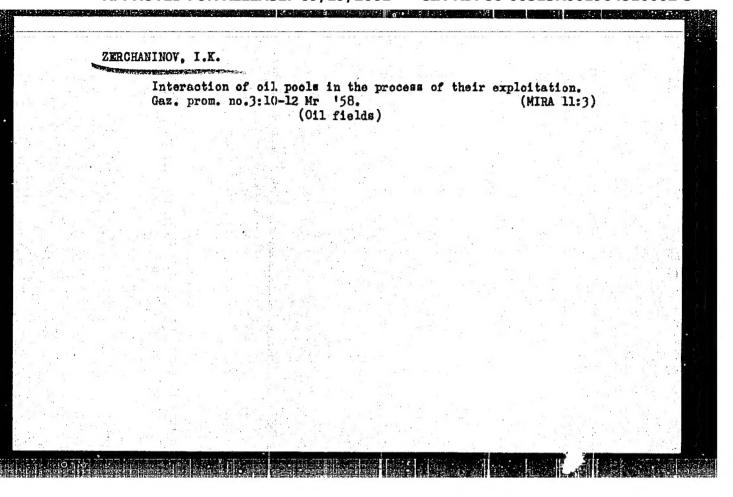
Mathod for studying water wells. Geol. nefti 1 gaza 5
no. 1:48-52 Ja '61.

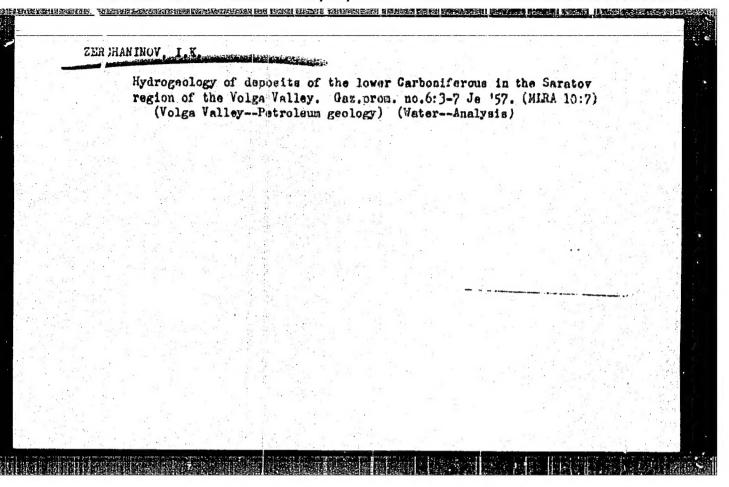
(MIRA 14:1)

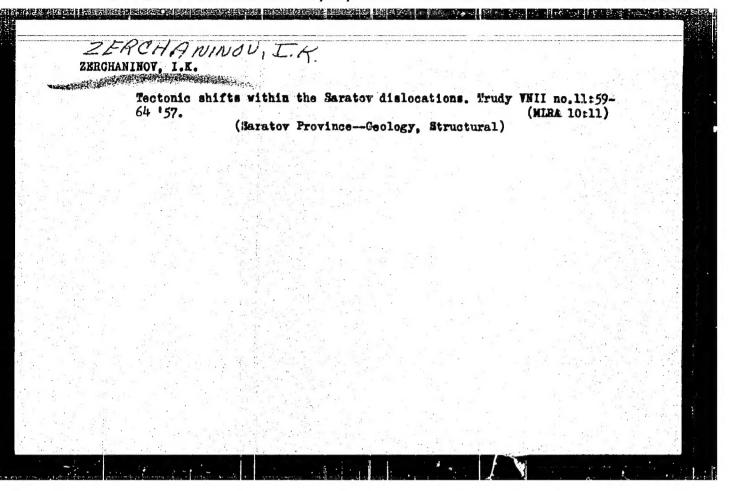
1. Vsesoyuznyy nefte-gazovyy nauchno-issledovatel'skiy institut. (Oil field brines)







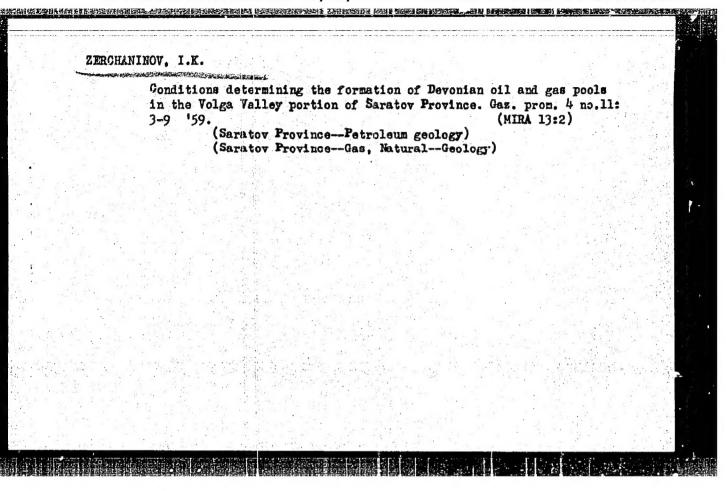


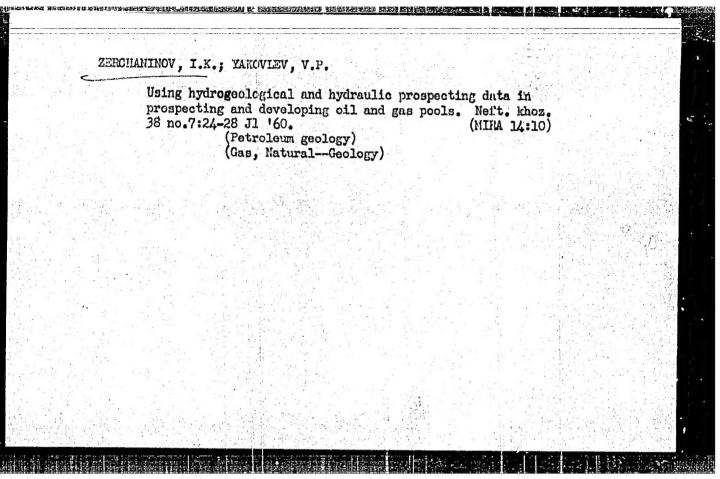


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"Formation of Terrace Structures of Second and Third Order in the Case of Saratov Displacements." All-Union Petroleum Gas Sci Res Inst (VNII) [Sic], Moscow, 1955. (Dissertation for the Degree of Candidate of Geological and Mineralogical Sciences.)

S0: M-972, 20 Feb 56





ZERCHANINOV, Igor' Konstantinovich; SHCHERBAKOV, G.V., red.; VORONOVA, V.V., tekhn. red.

[The technique of investigating water-tapping wells opening up productive deposits in the Volga-Ural region] Metodika issledovaniia vodianykh skvazhin, vskryvshikh produktivnye otlozheniia na territorii Uralo-Povolzhia. Moskva, Gostoptekhizdat, 1962. 77 p.

(Volga-Ural region-Oil field brines)

ZERCHANINOV, L. 'g.Vyazniki, Vladimirskoy oblasti)

Joyful voices of lettle children are ringing. Prom.koop.
14 no.7:36 Jl '60. (MIRA 13:8)

(Vyazniki--Kindergortens)

#### ZERCHANINOV, L.K.

Review of V.S.Miasoedov's monograph "Epidemiology of opisthorchosis." Med. paraz. i paraz. bol. 33 no.52627-628 S-0 '64. (MIRA 18:4)

USSR Zooparasitology - Parasitic Worms 195%, No. 66268 of. JOUR. RZBiol., Ro. 19 Zerchaninov, L.K., Sokolova, Yc.K. AUTHOR INST. : Opisthorchiasis and Diphyllobothriasis in the TITLE Everdlovskaya Oblast Ned. Parazitol, i Parazitarn. Bolezni, 1957, ORIG. PUB. : Vol.26, No.6, 714-717 In the northern regions of the Sverdlovskaya ABSTRACT Oblast, situated along the rivers Loz'va, Sos'va, and Tavda, foci were discovered of opisthorchiasis and diphyllobothriasis among persons and domestic animals. Up to 65% of merops from the reservoirs of the Taporinskiy Payon were infected with pleurocercoids. 1/1 CARD:

ZERCHININOV, L.K. Interprovince scientific and practical conference on the control of helminthiases in the regions of the Urals and Western Siberia and the scientific conference of the Tyumen' Branch of the Omsk Scientific Research Institute for Natural Focus Infections on medical parasitology. Med. paraz. i paraz. bol. 33 no.6:756

(MIRA 18:6) N-D 164.

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[Protozoan infections and helminthiases of the Ural population] Protozoinye bolezni i gel'mintozy naseleniia Urala. Moskva, Medgiz, 1961. 182 p. (MIRA 17:2)

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Bithymia leachi in the waters of Tyumen' Province. Med. paraz. i paraz. bol. 32 no.6:741 N-D '63 (MIRA 18:1)

l. Iz parazitologicheskoy laboratorii filiala Omskogo nauchno issledovatel skogo instituta prirodnochagovykh infektsiy v Tyumeni (ispolnyayushchiy obyazannosti direktora V.N. Shril ko).

PLOTNIKOV, N.N.; ZERCHANINOV, L.K.; YALDYGINA, Z.S.

Experimental treatment of opisthorchosis with hexachloro-p-xylene. Report No.2. Med.paraz.i paraz.bol. 33 no.4:387-392 Jl-Ag 64.

1. Klinicheskiy otdel Instituta meditsinskoy parazitologii i tropicheskoy meditsiny imeni Martsinovskogo i parazitologii i otdel filiala Omskogo instituta prirodnoochagovykh infektsiy v Tyumeni.

ZERCHANINOV, L.K.; SOKOLOVA, Ye.K.

Opisthorchiasis and diphyllobothriasis in Sverdlovez Province. Med. paras.i paras.bol. 26 no.6:714-717 N-D '57. (NIRA 13:4)

1. Is parasitologicheskogo otdela Sverdlovskogo nauchno-issledovatel'skogo instituta epidemiologii i mikrobiologii Ministerstva sdravookhraneniya RSFSR (direktor instituta G.F. Bogdanov). (SVEEDLOVSK PROVINCE-WORMS, INTESTIMAL AND PARASITIO) (LIVER FLUKE)

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1. Iz Sverdlovskogo instituta epidemiologii, mikrobiologii i gigiyeny (dir. instituta G.F.Bogdanov)

(ASCAPIASIS, epidemiol.

in Russia)

BOGDANOV, G.F., red.; BYCHKOVSKAYA, O.V., red.; ZERCHANINOV, L.K., red.; MEDVINSKAYA, K.G., red.; PERETTS, L.G., prof., red.; PUSHKAREVA, Z.V., red.; DAVYDOVA, I., red.; PAL'HINA, H., tekhn.red.

[Increasing the activity of antibiotics, sulfonamides, and blood serum; collection of articles] Uvelichenie aktivnosti antibiotikov, sulfamidov i krovianci syvorotki; sbornik statei. Sverdlovsk, 1957. 205 p. (MIRA 13:1)

1. Sverdlovskiy nauchno-issledovatel'skiy institut antibiotikov. (ANTIBIOTICS) (SULFONAMIDES) (SKRUM)

ZERCHANINOV, L.K.; KONDINSKIY, G.V.

Distribution of toxoplasmosis in Tyumen' Province. Zhur.mikrobiol.;
epid. 1 immun. 42 no.2:55-58 F '55. (MIRA 1816)

1. Filial Omskogo instituta prirodnoochagovykh infektsiy v Tyumeni.

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ACC NR: AP7009356

SOURCE CODE: GE/0064/66/018/11-/0431/0435

AUTHOR: Zerche, Manfred (Doctor)

ORG: Schwerin Weather Bureau, Schwerin (Amt fur Meteorologie Schwerin)

TITLE: Computation of estimated fog frequencies using 3-hr observations

SOURCE: Zeitschrift fur Meteorologie, v. 18, no. 11-12, 1966, 431-435

TOPIC TAGS: climatology, fog, meteorology, fog frequency, annual fog

frequency/North German plain

ABSTRACT: Annual and monthly fog frequencies were studied for the sake of climatology dealing with traffic conditions and health resort problems. Improving on Koppen's method and formula to calculate fog frequency in hours from long-term 3-hr observations, accurate determination of mean annual fog frequency variations (mean monthly fog frequencies in hours) in the North German inland plain by computation of annual occurrence estimates was shown. Koppen's method provided only an estimate of annual frequencies. Orig. art. has: 1 table and 6 formulas. [DR]

[Based on author's abstract]

SUB CODE: 04/SUBM DATE: none/ORIG REF: 002/OTH REF: 004/

UDC: 551, 575, 36

TALAKIN, O.G.; AKHANSHCHIKOVA, L.A.; SOSNOVSKIY, Ye.N.; PANKRATOV, A.V.;
ZERCHENINOV, A.N.

Heat of formation of fluonitrate. Zhur.fiz.khim. 36 no.5:1065(MIRA 15:8)

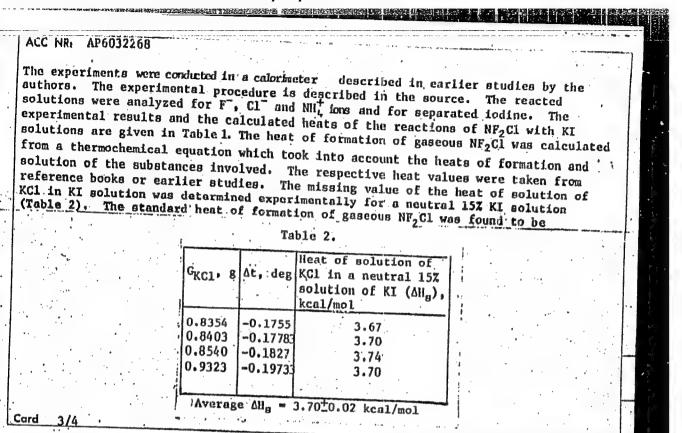
(Fluonitrate) (Heat of formation)

ACC NRI AP6032268 UR/0076/66/040/009/2101/2104 . SOURCE CODE: AUTHOR: Zercheninov, A. N.; Chesnokov, V. N.; Pankratov, A. V. ORG: none TITLE: Standard heat of formation of chlorodifluoramine SOURCE: Zhurnal fizicheakoy khimii, v. 40, no. 9, 2101-2104 TOPIC TAGS: chlorodifluoramine, heat of formation, potassium iodide solution, gaseous chlorodifluoramine, liquid chlorodifluoramine, FLUORINE COMPOUND ABSTRACT: The standard heat of formation of chlorodifluoramine has been determined from its reaction with an aqueous solution of potacsium iodide  $NF_2CI + I^- + H^+ \rightarrow I_2 + N_2F_4 + CI^- + F^- + NH_4^+ + N_2$ This reaction proceeds in several steps. Selection of proper [unspecified] pH of the solution and contact time of NF2Cl with the solution reduced reaction 1 to the reaction  $NF_2CI + \frac{3}{2}KI = KCI + \frac{1}{2}KI_3 + \frac{1}{2}N_2F_4$ (2) NF2Cl used in the experiments contained, in addition to N2 and N2C, 1 to 7% N2F2 whose presence caused in the calorimeter the additional reaction  $N_1P_1 + 3KI = 2KP + KI_2 + N_1$ (3)

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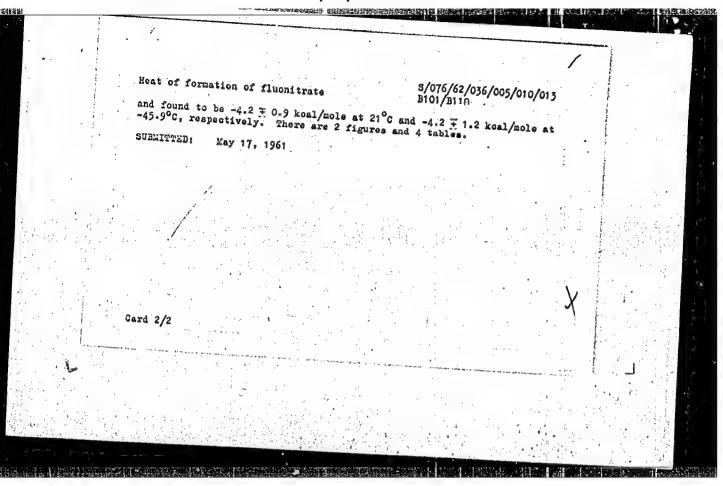
UDC: 541.11

ACC NRI AP6	032268		Table 1.			
	NF <sub>2</sub> Cl content in the sample,	Analysis of.	orned orned orned orned orned orned orned count teed teed teed	the calorineer (At), (alculated overall hear of reactions 2 and 3 (Qr), cal (Gr), cal reaction 3 (Q <sub>3</sub> ), caction 3 (Q <sub>3</sub> ),	reaction 2 (Q2), calculated heat of the reaction of NF2CI with KI solution (-AHz), kcal/mol	
	90, 1 83,6 83,6 83,6 83,6 83,6 83,6	0,420,0286 1,680 0,4010,0324 1,607 0,3720,0272 1,453 0,2613,178 1,096 0,2830,0131	0,40700,0091 0,6270, 0,98920,0464 1,614 1, 1,05090,0497 1,716 1, 0,98920,0563 1,652 1,052 1,054 1,513 1,054 1,513 1,054	6119 158,9 21,7 5363 473,4 107,3 5119 465,9 115,0 4496 446,7 130,2 297,1 399,8 109,4	137,2 29,48 366,1 32,37 350,9 29,20 316,5 27,68 244,7 33,20 244,2 30,60	
	76,6 76,6 62,1 36,1 36,1 36,1 36,1	0,179 0,0136	0,441610,0236, — 0, 0,7943]0,0162 1,215[1, 0,9275]0,0327 — 1, 1,3715[0,0396] 2,142[2, 0,7770]0,0268 1,236[1, 0,8214[0,0300] 1,307[1,	9327 212, 7 54, 6 4212 324, 0 37, 5 2577 379, 8 75, 6 7246 621, 1 91, 6 4973 341, 4 60, 6 4967 341, 2 69, 4 9242 301, 9 67, 1	158, 1 31, 31 286, 5 31, 55 304, 2 28, 69 529, 5 33, 77 274, 8 30, 93 271, 8 28, 94 244, 8 31, 11	



$2 \pm 2.9$ kcal/mol. The heat of formation of liquid NF <sub>2</sub> Cl was calculated by taking value of 4.35 kcal/mol for the heat of vaporization of NF <sub>2</sub> Cl at -67C (boiling), and in the assumption that the average heat capacity of NF <sub>2</sub> Cl in the radius of NF <sub>2</sub> Cl in the radius of NF <sub>2</sub> Cl at -67C was found to be -2.2 kcal/mol. The N-Cl bond energy was callated in the assumption that the NF <sub>2</sub> Cl at -87C was found to be -2.2 kcal/mol. The N-Cl bond energy was	iking
culated in the assumption that the N-F hand appear in NF of the	inge of
	in the
$E(N-Cl) = \Delta H_{f200}^{0}(Cl) + \Delta H_{f200}^{0}(NF_{2}) - \Delta H_{f200}(NF_{2}Cl) = 35.3 \text{ kcal/mol.}$	,
8. gree 1131. 3 fa0168.	
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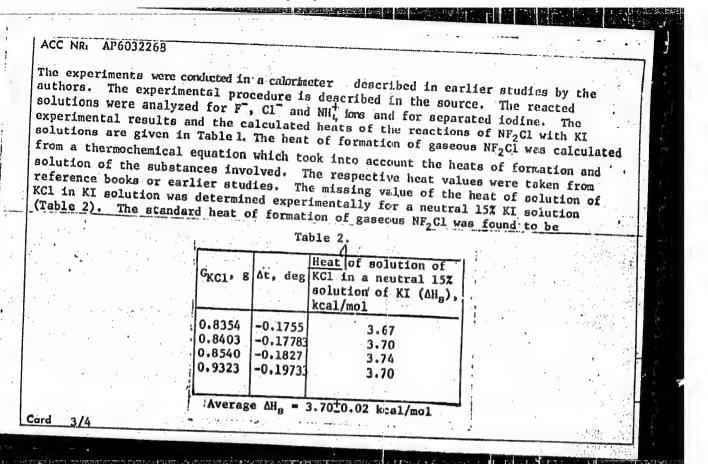
37635 s/076/62/036/005/010/013 B101/3110 11.2131 Talakin, O. G., Akhanshchikova, L. A., Sosnovskiy, Ye. N., AUTHORS: Pankratov, A. V., and Zercheninov, A. M. Heat of formation of fluonitrate TITLE: PERIODICAL: Zhurnal fizicheskoy khimii, v. 36, no. 5, 1962, 1065-1067 TEXT: The heat of formation of NO P was calorimetrically determined on the basis of the reaction NO3F + 2KOH - KNO3 + KF + 0.5 02 + H2O, the NC3F being synthosized by bubbling  $F_2$  through  $HNO_3$  thus:  $HNO_3 + F_2 = HP + NO_3P$ . The HP was absorbed by KF, and NO,F was condensed at -1830C. The heats (keal/mole) of reaction between NO3F and KOH (Q = 93.5 7 0.8), between KP and KOH ( $Q_3 = 3.35 \mp 0.011$ ), and between KNO<sub>3</sub> and KOH ( $Q_4 = -5.93 \mp 0.023$ ) were measured with a calorimeter calibrated with KCl. From the system of equations which allows for this and the other side reactions of the process the heats of formation of gaseous and liquid NO<sub>3</sub>P were calculated

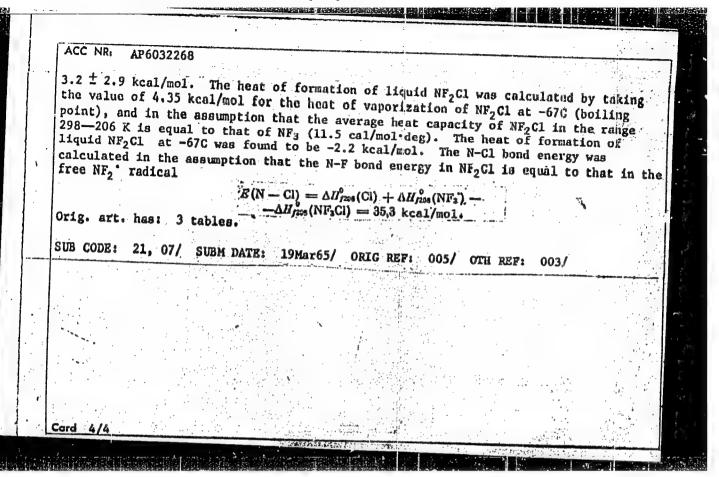


ACC NR. AP6032268 SOURCE CODE: UR/0076/66/040/009/2101/2104 AUTHOR: Zercheninov, A. N.; Chesnokov, V. N.; Pankratov, A. V. ORG: none TITLE: Standard heat of formation of chlorodifluoramine SOURCE: Zhurnal fizicheskoy khimii, v. 40, no. 9, 2101-2104 TOPIC TAGS: chlorodifluoramine, heat of formation, potassium iodide solution, gaseous chlorodifluoramine, liquid chlorodifluoramine, FLUORINE COMPOUND, ABSTRACT: The standard heat of formation of chlorodifluoramine has been determined from its reaction with an aqueous solution of potassium iodide  $NF_2CI + I - + H + + I_2 + N_2F_4 + CI - + F - + NH_4 + + N_2$ This reaction proceeds in eseveral steps. Selection of proper [unspecified] pH of the solution and contact time of NF2Cl with the solution reduced reaction 1 to the  $NF_2CI + \frac{1}{2}KI = KCI + \frac{1}{2}KI_3 + \frac{1}{2}N_2F_6$ (2)  $NF_2C1$  used in the experiments contained, in addition to  $N_2$  and  $N_2O_2$  1 to 7%  $N_2F_2$ whose presence caused in the calorimeter the additional reaction  $N_s F_s + 3KI = 2KF + KI_s + N_s$ 1/4 Card (3)UDC:

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		<del> </del>		Table	1.				
	NF2G1	Analys	is of th	e Bolut	ion	ne tn		2 원년	-
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		279 0,0142	1917,08		1,03311	32421 301,9	69,4 271 57,1 244 5±1.6 kc	8 31 11	





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ACCESSION NR: AP300	2942	s/0076/63/037/006/1399/1401 7/
AUTHOR: Fankratov,	A. V.; Zercheninov, A. N.;	Milekin, O. G.; Sokolov, C. M.;
TITLE: Standard ent	calpy of formation of the a	whive isomer of diffuorodiszine
The second of the second secon	icheskoy khimii, v. 37, no.	
经产品收益 化二甲基乙二甲基甲基二甲二二甲基二甲二甲二甲基甲二甲	为我们,我们是我们的时间,这一点,对此一个人,一个人的,你就是我们的一定,我就会一个人都是这个是这个	Ufluorodiazine, IR measurement
ABSTRACT: The stends	rd enthelmy for the govern-	
ABSTRACT: The stend was calculated. It for the liquid at 10	ord enthalpy for the gaseous vas 25.3 + or - 2.0 kcal/mol 05.7°. Data was obtained by the lacer with an acid relation	active isomer of difluorodissins.  It was 20.5 + or = 2.0 kcal/mil  IR measurement of the heat of ion of KI. Orig. art. has: 2
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ABSTRACT: The stends was calculated. It for the liquid at -10 reaction of the activitables, 1 figure, and ASSOCIATION: none	erd enthelpy for the gaseous was 25.3 + or - 2.0 kcal/mol 05.7°. Data was obtained by we isomer with en acid solut 1.3 equations.	active isomer of diffuorodissing It was 20.5 for = 2.0 kcal/mil IR measurement of the heat of ion of KL. Orig. art. has: 2
ABSTRACT: The stender was calculated. It is for the liquid at all reaction of the activitables, 1 figure, and ASSOCIATION: none SUBMITTED: 22Aug62	prd enthelpy for the gaseous vas 25.3 + or - 2.0 kcal/mol 05.7°. Data was obtained by re isomer with en acid solut 1.3 equations.  DATE ACQ: 16Jul63	active isomer of diffuorodissins.  It was 20.5 for - 2.0 kcal/mil.  IR measurement of the heat of ion of KI. Orig. art. has: 2

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Determining the smoothness of yarns and threads. p. 773. (GLASNIK, Vol. 6, No. 9, Sept. 1957

SO: Monthly List of East European Accessions (EEAL) LC Vol. 6, Nol 12, Dec. 1957

ZERDIK, M.

Textile-research institutions in Sweden and Denmark. p. 42.

Periodical: TEKSTIL

Vol. 8, no. 1, Jan. 1959.

TECHNOLOGY

SO: Monthly List of East European Accessions (EEAI) LC

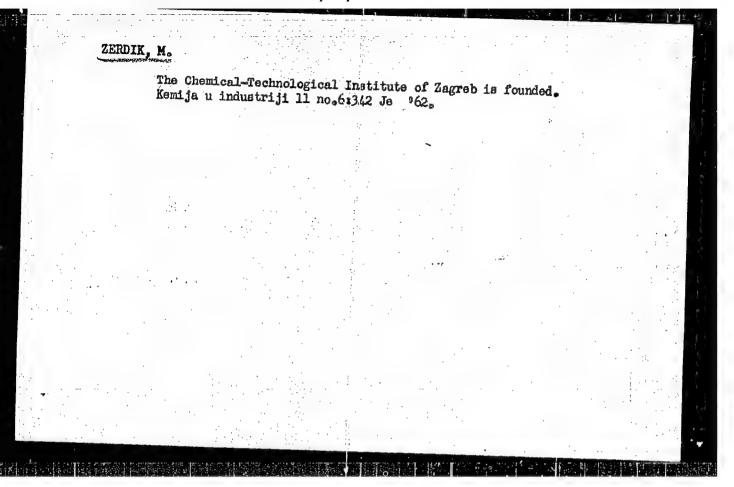
Vol. 8, no. 4 April 1959, Uncl.

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Charlest damage and swelling of cotton fibers in strong alkalies. Pt. 2. Tekstil Eagreb 13 no.10:243-852 0 164.

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Starch adhesives in the textile industry. Kem ind 12 no.5:

1. Faculty of Technology, University of Zagreb, Zagreb.

ZERDIK, M.

Scientific progress and modern chemical technology of textiles. p. 729
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Vol. 5, no. 9, September 1956

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Yugoslavia (430)

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A rapid method for the determination of grease in wool. p. 293. TEHNICKI PREGLED. (Croatia. Uprava za unapredenje proizvodnje pri privednom savjetu) Zagreb. (Bimonthly technical journal insued by the Production Improvement Administration of the Economic Council) No. 6, 1951.

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ZERDIK, Mladen, prof., dipl. inz.; RAFFAELLI, Dubravka, dipl. inz., asistent

Self-inflammability of oiled raw silk. Tekstil Zagreb 18 no. 1: 14-20 Ja '64.

- Predstojnik Zavoda za tekstilnu kemijsku tehnologiju Tehnoloskog fakulteta Sveuculista u Zagrebu (for Zerdik).
- 2. Zavod za tekstilnu kemijsku tehnologiju Tehnoloskog fakulteta Sveucilista u Zagrebu (for Raffaelli).

"Power apparatus for cutting and saving wood" p. 20 (las polski, Vol. 26, No. 2, Feb. 1952, Warszawa)

East European Vol. 3, No. 3

So: Monthly List of RAMMAN Accessions Library of Congress, March 1954, Uncl.

ZEREBECKI, JAN.

"Mechanizacja pracy przy scince i wyrobce drewna (1. wyd.) Warszawa, Panstwowe Wydawn. Rolnicize i Lesne, 1951. 59 p. (Biblioteczka lesna) (Mechanization of the cutting of timber and milling. lst ed.)

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Not In DLC

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ZEREBECKE, J.

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SO: Monthly List of East European Accessions, L. C., Vol. 3, No. 4, April, 1954

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LEREBISOVA, K.I.

SUBJECT

USSR / PHYSICS

CARD 1 / 2

PA - 1406

AUTHOR TITLE

NEMILOV, JU.A., ZEREBCOVA, K.I., FUNSTEJN, B.L.

On the Relationship between the Processes of Stripping and the

Production of a Compound Nucleus on the Occasion of Reaction with

Deuterons.

PERIODICAL

Žurn.eksp.i teor.fis, 30, fasc.6, 1013-1016 (1956)

Issued: 8 / 1956 reviewed: 10 / 1956

The relationship of these reactions on the nucleus  $Mg^{26}(d,p)Mg^{27}$  is here estimated by comparison of the yields of those nuclei which are produced on the occasion of d,p-processes and d,a-processes on an  $Mg^{26}$  nucleus as well as on the occasion of n,p-processes and n,a-processes on an  $Mg^{26}$  nucleus. For the purpose of a more accurate quantitative estimation of the relative probabilities of the two mechanisms mentioned in the title two reactions are selected (one of them with deuterons and the other with any other data as e.g. neutrons), in which one and the same compound nucleus is produced. The radioactive nuclei  $B_1$  and  $B_2$  created in connection with the reactions selected on this occasion had decay periods which, from the point of view of measuring technique, were favorable. The ratios of the quantities of radioactive nuclei  $B_1$  and  $B_2$  produced in the targets by irradiation with deuterons and neutrons were determined from the fading curves of radioactivity. It is true that:  $\sigma(d,p)/\sigma(d,a)=(\sigma(d,P)_{c.n.}+\sigma(d,p))$  strip  $\varphi F)/\sigma(d,a)=N_1$  and  $\sigma(n,p)/\sigma(n,a)=N_2$ . Here cn. refers to a compound nucleus, strip to a stripping process, and F denotes the term due to the interference between the two terms. As the decay of the compound

Zurn.eksp.i teor.fis, 30, fasc.6, 1013-1016 (1956) CARD 2 / 2 PA - 1406 nucleus does not depend on the manner in which it was produced, it is true that  $\sigma(d,p)_{c.n.}/\sigma(d,\alpha)=\sigma(n,p)/\sigma(n,\alpha)$  and herefrom follows

' $(\sigma(d,p)\text{strip}^+F)/\sigma(d,p)_{c.n.} = (N_1-N_2)/N_2$ . This is correct only if the compound nuclei produced by the capture of a deuteron and of a neutron have the same excitation energies. However, also if these conditions are satisfied with accuracy, it is possible that inaccuracies occur as a result of the influence exercised by broad overlapping levels. No resonance phenomena were, by the way, found. The following two reactions were selected in this case:

The ratios of radioactivities found are represented in diagrams as functions of the deuteron energy. In order to obtain neutrons with uniform energies the reaction  $D^2 + D^2$  was used. As a deuteron target a circonium layer irradiated for a

action D + D was used. As a deuteron target a circonium layer irradiated for a long time with slow deuterons(0,8 MeV) was used.

On the occasion of reaction with deuterons and reaction with neutrons the proton

On the occasion of reaction with deuterons and reaction with neutrons the proton yield increases as against the yield of  $\alpha$ -particles as a result of a decrease of deuteron energy. The ratio  $(\sigma_{\text{strip}}^+ + F)/\sigma_{\text{c.n.}}$  has maximum deuteron energies of from

1 to 2 MeV and a value of 8 to 9.

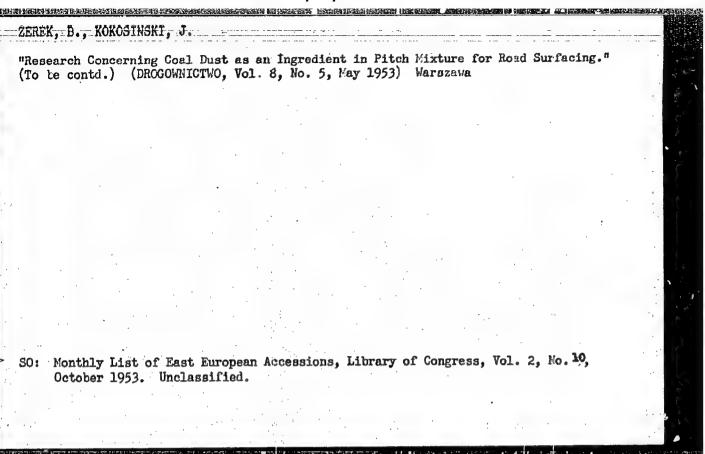
INSTITUTION: Radium Institute of the Academy of Science

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(Electric locomotives)

ident as director 64

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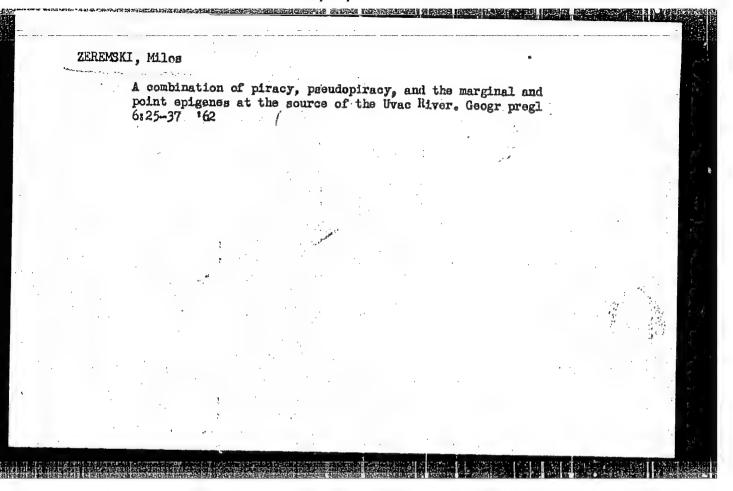
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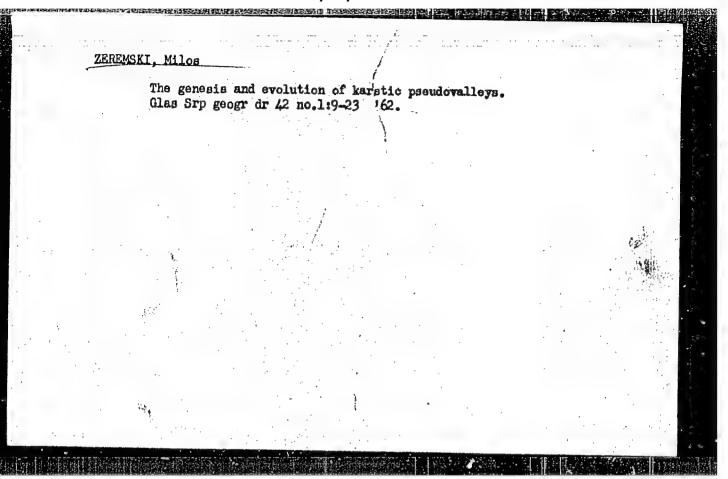
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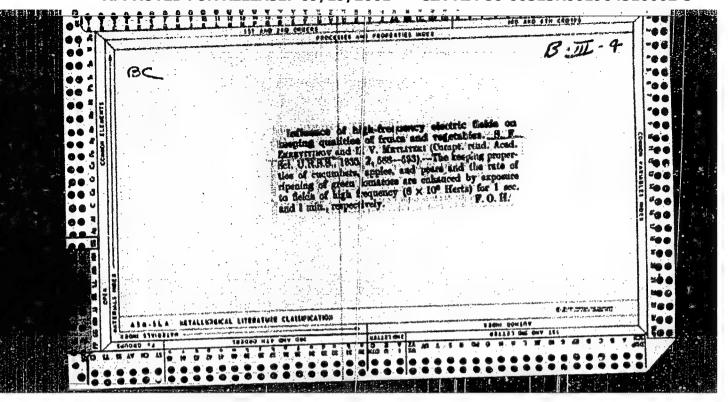
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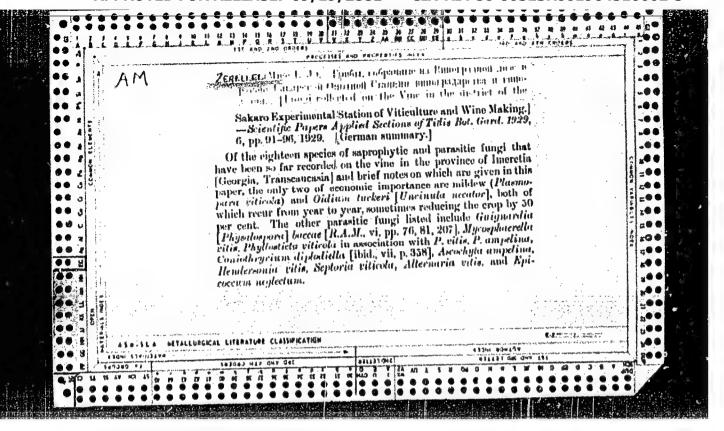
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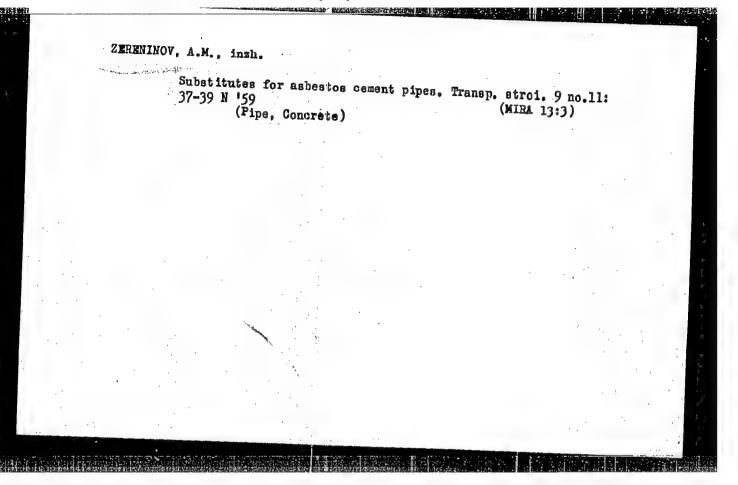
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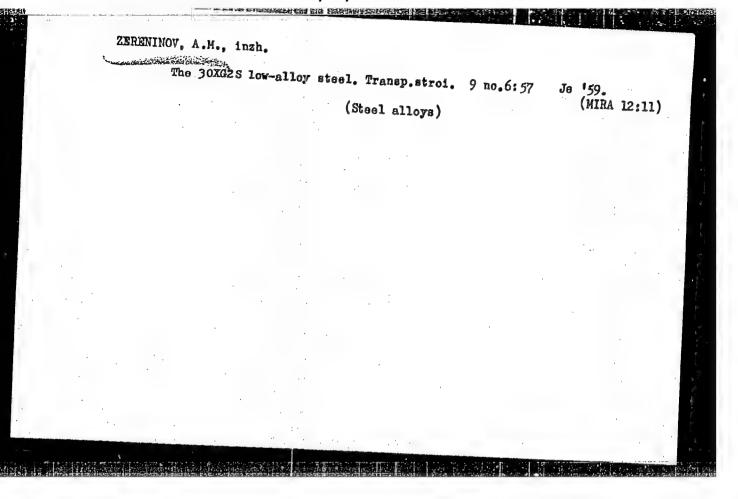
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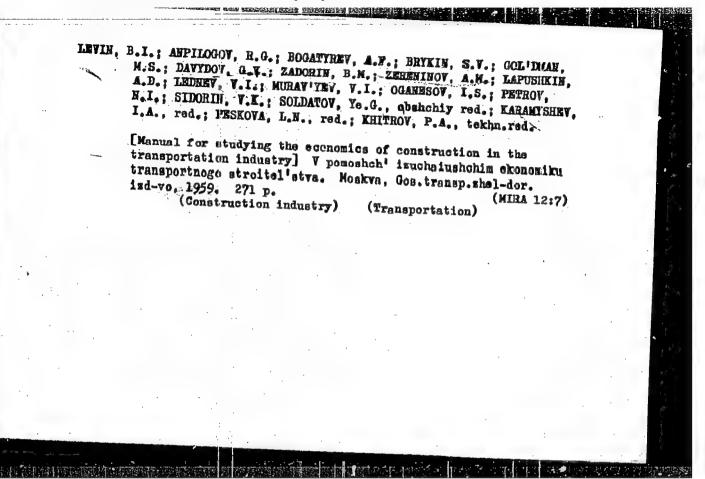
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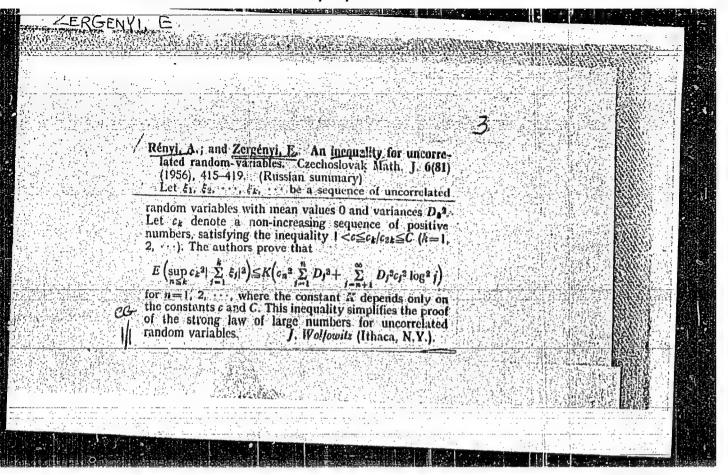
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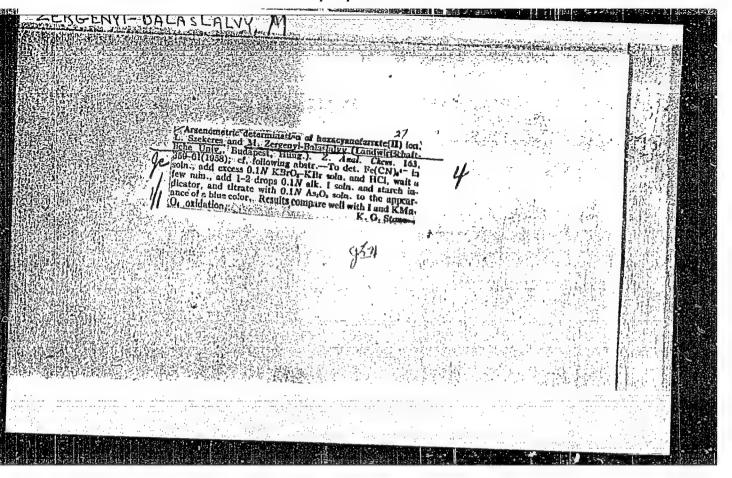
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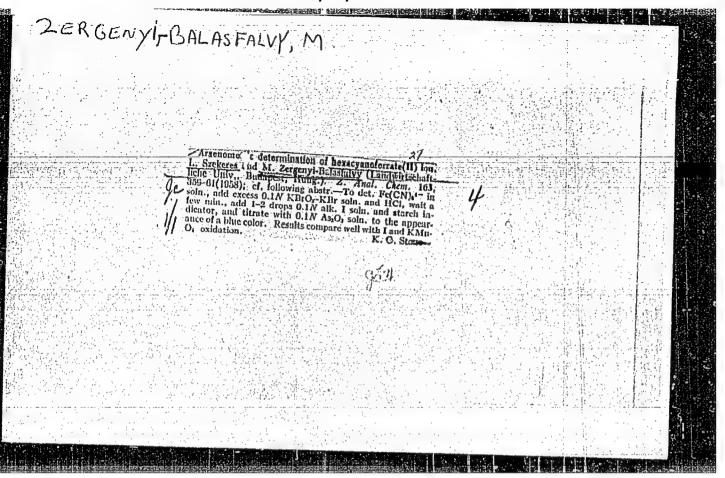
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Abstract: An arsenometric method of oxalate ion determination based on the exidation of exalate ions with bromine and the titration of the excessive bromine with arsenite solution was developed. 10 ml of 2 m. HCl is added to 10 to 20 ml of 0.1 m. bromidebromate solution containing 15 to 20 g of KDr per liter, after which 7.5 ml of 5 n. NaOH solution, the oxalate solution under study (about 5 to 10 ml of 0.1 n

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ZERGOLLERII. Josio: and ZERGOLLERII, Sergije; Department of Orthopedics and Institute of Radiology (Ortopedaki odjel i Zavod za radiologiju bolnice) "Dr. Josip Kajfes", Zagrab

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